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The Mole Supplemental Problems KEY. 1. Identify and calculate the number of representative particles in each of the following quantities. a. 2.15 moles of gold. b. 0.151 mole of nitrogen oxide . c. 11.5 moles of potassium bromide . 2. Calculate the number of moles of the substance that contains the following number of representative particles.

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level of water rises to 18.0 mL.

Aluminum has a density of 2.7 g/mL.

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308 Chapter 11 The Mole CHAPTER
11 What You'll Learn You will use
the mole and molar mass to make
conver-sions among moles, mass,

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and number of representative particles. You will determine the percent composition of the components of compounds. You will calculate the empirical and molecular formulas for compounds and determine the formulas for hydrates.

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310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02×10^{23} or one mole of representative particles. The representative particle for each

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substance is shown in a box. Refer to Table C-1 in Appendix C for a key to atom color conventions. Molecule Atom Formula unit H 2O Cu NaCl

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*Supplemental Problems Problem 4
Godwit Associates paid \$60,000 for a 20-seat skybox at Memorial Stadium for eight professional football games. Regular seats to these games range from \$70 to \$150 each. At one game, an employee of*

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Godwit entertained 18 clients.

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Supplemental Problems Chemistry: Matter and Change • Chapter 2 1 Data Analysis Data Analysis 1. A

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sample of aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample. 2. Saturn is about 1 429 000 km from the Sun.

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comprehensive discussion of these

disorders, Mental Disorders and

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Children provides the best currently

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available information regarding demographics, diagnosis, treatment, and expectations for the disorder time course - both

Chapter 11 Supplemental Problems Stoichiometry Answers

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Stoichiometry Answers Chapter 11 Supplemental Problems The Mole Answer Key We can see from the stoichiometry of the reaction that $3/2$ mol of O_2 is required to produce 1 mol of H_2SO_4 . This is a standard stoichiometry problem of the type presented in Section 11.4, except this problem asks for the volume Page 28/31

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this set (32) one mole is = to how many particles. 6.02×10 to the 23rd. abbreviation for a mole is. mol. mole is the number of representative carbon 12 particles in how many grams of carbon 12.

Chapter 11: The Mole - Neshaminy

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310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02×10^{23} or one mole of representative particles. The representative particle for each substance is shown in a box. Refer to Table C-1 in Appendix C for a key to atom color conventions. Molecule Atom Formula unit H₂O Cu NaCl

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this set (32) one mole is = to how
many particles. 6.02×10 to the 23rd.
abbreviation for a mole is. mol. mole
is the number of representative
carbon 12 particles in how many*

grams of carbon 12.

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Data Analysis 1. A sample of aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample. 2. Saturn is about 1 429 000 km from the Sun.

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CHAPTER 10 SOLUTIONS

MANUAL 10. Explain how a mole is similar to a dozen. The mole is a unit for counting 6.02×10^{23}

representative particles. The dozen is used to count 12 items. 11. Apply

How does a chemist count the number of particles in a given number of moles of a substance?

They multiply the number moles by Avogadro's number. 12.

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1. 1 mole = 6.02×10^{23} atoms 2.

No, your friend is wrong. The units you will end up with are (atoms) $2 / \text{mole}$, which is not what you want.

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CHAPTER 10 SOLUTIONS

MANUAL 10. Explain how a mole is similar to a dozen. The mole is a unit for counting 6.02×10^{23}

representative particles. The dozen is used to count 12 items. 11. Apply

How does a chemist count the number of particles in a given

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number of moles of a substance?

*They multiply the number moles by
Avogadro's number. 12.*

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Stoichiometry

CHAPTER 11 SOLUTIONS

MANUAL Section 11.1 Defining Stoichiometry pages 368–372

Practice Problems pages 371–372 1.

Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed. a. $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ 1 molecule N_2 3 molecules H_2 ? 2 molecules NH_3 ...

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Chapter 11. 1. Silicon nitride is used in the manufacturing of high-temperature thermal insulation for heat engines and turbines. It is produced by the following reaction.

$$3\text{Si}(s) + 2\text{N}_2(g) \rightarrow \text{Si}_3\text{N}_4(s)$$

a. Interpret the equation in terms of particles, moles, and masses. Particles: 3 silicon atoms 2 nitrogen molecules 1 formula unit silicon nitride

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Section 10.1 Measuring Matter 1.

pair 2. 5 3. dozen 4. gross 5. 200 6.

ream 7. 6,000,000,000 8. 0.5 mol 9.

6.02 10²³ 10. four moles 11. 6.02 10

Cu atoms 23 1 mol Cu 12. 4 23 4 1

mol CH 6.02 10 molecules CH 13. 23

1 mol Xe

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answers**

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Avogadro's Number Review
Questions 1. The mole is an important measurement in chemistry.
Chapter 10 The Mole Worksheet
Answers The mole is an important concept for talking about a very large number of Page 13/28.*

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