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### Conceptual Questions

23. For the chemical equilibrium  $A + 2B \rightleftharpoons 2C$ , the value of the equilibrium constant,  $K$ , is 10. What is the value of the equilibrium constant for the following reaction?  $2A + 4B \rightleftharpoons 4C$   $K =$  ??? Given that:  $A + 2B \rightleftharpoons 2C$   $K = 10$  a. 0.10 d. 400 b. 0.20 e. 100 c. 40

24. For the chemical equilibrium  $A + 2B \rightleftharpoons 2C$ , the value of the equilibrium constant,  $K$ , is 10

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*examples of problems where  
equilibrium calculations are useful.  
Examples: · Diatoms exist in surface  
seawater. The solubility of diatom  
shell ....*

**Chapter 16 Chemical Equilibrium  
Solutions to Practice ...**

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*Chapter 16 Chemical Equilibrium  
Solutions to Practice Problems 1. ...  
Write the chemical equation for the  
equilibrium system. Write the  
equilibrium expression and substitute  
the equilibrium molar concentrations  
into the expression for  $K_c$ . Calculate  
the value of  $K_c$ . Act on Your Strategy*

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*Chemical Equilibrium Exam1 and  
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reaction is in equilibrium;  $X(g) + 2Y(g) \rightleftharpoons Z(g) + H$*

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*Chemical Equilibrium. For each of the following reactions, use Figure 16. Part 3 - sample questions. If a gaseous/aqueous reactant or product is added to a system at equilibrium, the system will ...*

**Lecture 9 - Chemical Equilibrium  
and Solution Chemistry ...**

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*1. Is a geochemical system at chemical equilibrium? 2. If not, what reaction(s) are most likely to occur? 3. What are the relative amounts of energy and materials that might occur in a system? Here are some examples of problems where equilibrium calculations are ...*

### **Equilibrium Constant - Practice Problems for Assignment 5**

*Determine the value of the equilibrium constant,  $K_c$ , for the reaction. Initially, a mixture of 0.100 M NO, 0.050 M H<sub>2</sub>, 0.100 M H<sub>2</sub>O was allowed to reach equilibrium (initially there was no N<sub>2</sub>). At*



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*equilibrium the concentration of NO was found to be 0.062 M. 9. Consider the following reaction  $N_2O_4 (g) \rightleftharpoons 2 NO_2 (g)$*

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*2004 Free Response – Form B 1.*

*$N_2(g) + 3 H_2(g) \rightleftharpoons 2 NH_3(g)$  For the  
reaction represented above, the  
value of the equilibrium constant,  $K_p$*

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is  $3.1 \times 10^{-4}$  at 700 K. a) Write the expression for the equilibrium constant,  $K_p$ , for the reaction. b) Assume that the initial partial pressures of the gases are as follows:

**Equilibrium Constant - Practice Problems for Assignment 5**

Determine the value of the equilibrium constant,  $K_c$ , for the reaction. Initially, a mixture of 0.100 M NO, 0.050 M  $H_2$ , 0.100 M  $H_2O$  was allowed to reach equilibrium (initially there was no  $N_2$ ). At equilibrium the concentration of NO was found to be 0.062 M. 9. Consider the following reaction  $N_2O_4(g) \rightleftharpoons 2$

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*NO<sub>2</sub> (g)*

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*1. Is a geochemical system at  
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*pressures of the gases are as follows:*

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*chemical is raised to the power of its coefficient in the chemical equation.*

*Solution for Practice Problems*

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*Problem The following reaction took place in a sealed flask at 250°C. PCl*

*5(g) PCl 3(g) + Cl 2(g) At*

*equilibrium, the gases in the flask*

*had the following concentrations:*

*[PCl 5] =  $1.2 \times 10^{-2}$  mol/L, [PCl*

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 $2 \text{NO} (g) \rightleftharpoons \text{N}_2 (g) + \text{O}_2 (g)$   
 $K_c = 7.4 \times 10^{-6}$  at 25 C From this we can conclude: (a) The reaction is at equilibrium.

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