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*1. The rate of a chemical reaction can be expressed in .
a. energy released per mole of reactant . b. grams per mole of reactant
c. moles per liter of solution d. volume of gas per minute*

2. Rate constant . a. is the proportionality constant in the rate law

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Sample Problem 14.2 SOLUTION: Determining Reaction Orders from Rate Laws PLAN: We inspect the exponents in the rate law, not the coefficients of the balanced equation, to find the individual orders. We add the individual orders to get the overall reaction order. (a) The exponent of $[NO]$ is 2 and the exponent of $[O_2]$ is 1, so the

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Chemical Reactions Exam1 and Problem Solutions. 1. Balance following chemical reaction; $C_2H_5OH + O_2 \rightarrow CO_2 + H_2O$. Solution: $C_2H_5OH + O_2 \rightarrow CO_2 + H_2O$. We have 2 C atoms in reactants but 1 in products. So, we multiply CO_2 with 2 to balance number of C atoms. $C_2H_5OH + O_2 \rightarrow 2CO_2 + H_2O$.

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pdf where M is any molecule, including cyclopropane.

Only those cyclopropane molecules with sufficient energy (denoted with an asterisk) can rearrange to propylene.

Which step determines the rate constant of the overall reaction?

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