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Molecular and Ionic Compounds – Chemistry

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These molecular compounds (covalent compounds) result when atoms share, rather than transfer (gain or lose), electrons. Covalent bonding is an important and extensive concept in chemistry, and it will be treated in considerable detail in a later chapter of this text. We can often identify molecular compounds on the basis of their physical ...

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The Kinetic-Molecular Theory – Chemistry

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Key Concepts and Summary The kinetic molecular theory is a simple but very effective model that effectively explains ideal gas behavior. The theory assumes that gases consist of widely separated molecules of negligible volume that are in constant motion, colliding elastically with one another and the walls of their container with average velocities determined by their absolute

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temperatures.

Ch. 1 Introduction - Chemistry 2e | OpenStax

Figure 1.1 Chemical substances and processes are essential for our existence, providing sustenance, keeping us clean and healthy, fabricating electronic devices, enabling transportation, and much more. (credit “left”: modification of work by “vxla”/Flickr; credit “left middle”: modification of work by “the Italian voice”/Flickr; credit “right middle”: modification of

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Isoborneol | C₁₀H₁₈O - PubChem

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Borneol is consumed excessively in China and Southeast Asian countries particularly in combined formula for preventing cardiovascular disease, but few studies were conducted on its effects on thrombosis. In this study, the antithrombotic and antiplatelet activities of borneol were investigated on thrombosis in vivo and on platelet aggregation ex-vivo.

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The Ionization of Hydrated Metal Ions. Unlike the group 1 and 2 metal ions of the preceding examples (Na^+ , Ca^{2+} , etc.), some metal ions function as acids in aqueous solutions. These ions are not just loosely solvated by water molecules when dissolved, instead they are covalently bonded to a fixed number of water molecules to yield a complex ion (see chapter on coordination

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chemistry).

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Compounds that have identical molecular formulas but differ in the bonding sequence of the atoms are called structural isomers. The monosaccharides, glucose, galactose, and fructose all have the same molecular formula, $C_6H_{12}O_6$, but we can see from Figure 3 that the ...

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3.1 Formula Mass and the Mole Concept – Chemistry

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Key Concepts and Summary. The formula mass of a substance is the sum of the average atomic masses of each atom represented in the chemical formula and is expressed in atomic mass units. The formula mass of a covalent compound is also called the molecular mass. A convenient amount unit for expressing very large numbers of atoms or molecules is ...

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7 Chemical Formulas and Chemical Compounds

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Chemical Formulas and Chemical Compounds SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Write empirical formulas to match the following molecular formulas: CH_3O_2 a. $C_2H_6O_4$ N_2O_5 b. N_2O_5 $HgCl$ c. Hg_2Cl_2 CH_2 d. C_6H_{12} 2. C_4H_8 A certain hydrocarbon has an empirical formula of CH_2 and a molar mass of ...

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